

## Mass And Mole Relationships In A Chemical Reaction Lab Answers

atomic mass is the concept related to a single atom whereas molecular mass relates to a group of atoms if you are planning to proceed with this concept then it is necessary to gain precise

mole can be defined as the amount of substance that contains the same number of entities atoms molecules ions etc as the number of atoms present in 12g of the carbon 12 isotope

define mole fraction a solution of sucrose in water is labelled as 20 w/w what would be the mole fraction of each component in the solution

consider the reaction  $2\text{SO}_2 + \text{O}_2 \rightleftharpoons 2\text{SO}_3$  for which  $K_c = 278\text{M}^{-1}$ . If 0.001 mole of each of the reagents  $\text{SO}_2$ ,  $\text{O}_2$  and  $\text{SO}_3$  are mixed in a 1.0 l flask determine the reaction quotient of the

a triatomic linear neglect vibration degree of freedom gas of two mole is taken through a reversible process ideal starting from a as shown in the figure the volume

ratio  $\frac{dfrac{v}{b}}{v}$

find the average translational kinetic energy per molecule if one mole of the gas is contained in a volume  $1.23 \times 10^{-3} \text{ m}^3$  at a pressure  $2.105 \text{ N m}^{-2}$  avagadro s number is  $6.02 \times 10^{23}$  molecules moles

click here to get an answer to your question 42 but no lol vessel and allowed to reach  $24.3 \times 10^{-3} \text{ mole of H}_2$

the compressibility factor for 1 mole of a vander waals gas at  $273 \text{ K}$   $8.21 \text{ atm}$  pressure having volume  $1.092 \text{ litre}$  is  $0.4$  assuming the volume of gas molecules to be negligible

the vapour pressure of two volatile liquid mixtures is  $P = 5.32 \times x_B$  in mm of hg where  $x_B$  is mole fraction of  $B$  in mixture what is the ratio of mole fractions of  $A$   $B$  in liquid phase

a solution of 5 g haemoglobin in 100 ml solution shows a temperature rise of  $0.031 \text{ C}$  for complete oxygenation each mole of haemoglobin binds 4 mole of oxygen if the heat capacity of the solution

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